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Calculate Percent Yield with Ideal Stoichiometry - Practice - 1 When 50.0 g of silicon dioxide is heated with an excess of carbon, 32.2 g of silicon carbide is produced. $\text{SiO}_2\text{(s)} + 3\text{C(s)} \rightarrow \text{SiC(s)} + \dots$

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Calculate Percent Yield with Ideal Stoichiometry - Practice - 2 If the reaction below proceeds with a 96.8% yield, how many kilograms of CaSO_4 are formed when 5.24 kg SO_2 reacts with an ...

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